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6. 2 classifying the elements

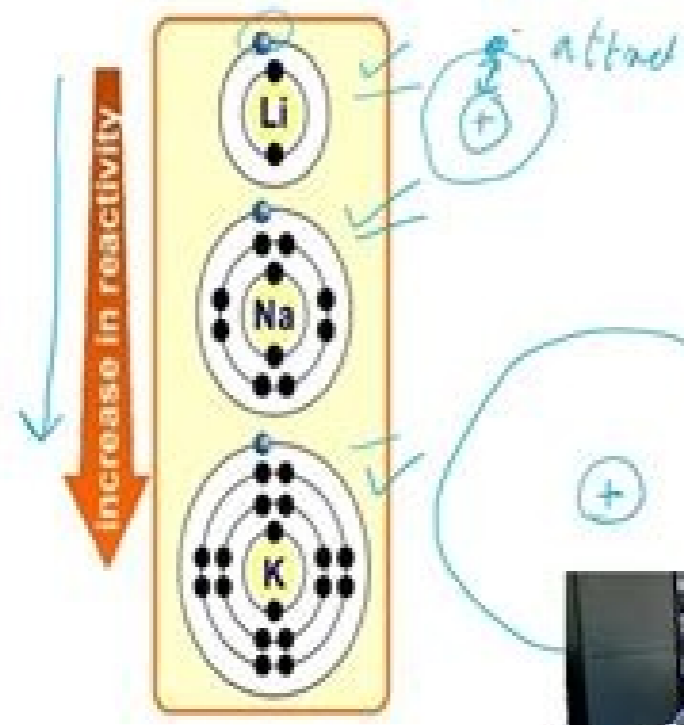
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Periodic Table of Elements

Sn, Xe, Rb, Sr
Xe
Rn, He, Xe, Kr
Pb, Ba, Cs, At

Group 1 reactivity

All the group 1 metals are reactive, but they get more reactive as you go down the group. This can be explained by looking at the electronic structure of the atoms:



Classification of elements

- | metals | non-metals |
|--|------------------|
| 1) Solids | 1) Liquids/Gases |
| 2) High density | 2) Least density |
| 3) malleable
← sheets | 3) Non-malleable |
| 4) ductile
← gold, silver, Al, Cu, Fe, Ni, etc. | 4) Non-lustrous |
| | 5) Non-ductile |

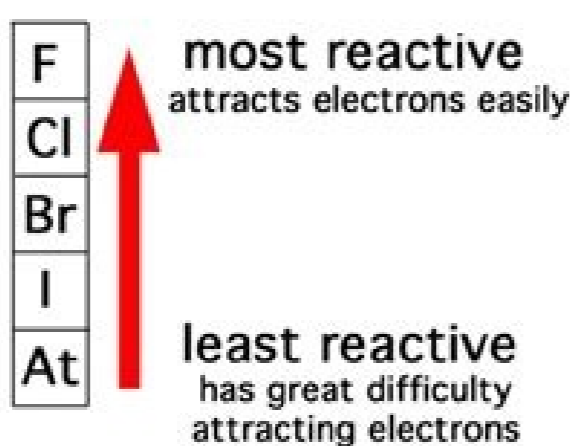
GCSE Science Daily Revision Task



Task: Produce a flash card on the elements in Group 7

Things to include:

The group 7 elements are also known as the halogens. They include fluorine, chlorine, bromine and iodine, which all have seven electrons in their outer shell. In a displacement reaction, a less reactive element is displaced by a more reactive element.



6.2 classifying the elements answers. 6.2 classifying the elements quizlet. 6.2 classifying the elements worksheet answers. 6.2 classifying the elements. 6.2 classifying the elements answer key.

16 Electron configurations in elements of the group 17 elements that appear below the main body of the periodic table. Next is 4S² and 3D⁸. 12 Electron configurations in the representative elements of the s and p sublevels of the busy energy level higher are not full. 14 Electron configurations in maneuvering elements in the periodic table, the elements of group B separate groups A on the left side of the table from groups A on the right side. For any representative element, its group number is equal to the number of electrons at the highest occupied energy level. There are two types of transition elements: d and f. The period indicates the highest level of main energy occupied. 1 Chapter 6 The periodic table 6.2 Classification of elements 6.1 Organization of elements 6.2 Classification of elements 6.3 Periodic trends Copyright © Pearson Education, Inc., or its affiliates. They are classified according to their electron configurations. The elements in group 1a are called alkali metals. 8 Electron configurations in groups 1a through 7a are often referred to as representative elements because they display a wide range of both physical and chemical properties. 15 Electron configurations in elements of the group Transition metals are the elements of group B that are generally shown in the main body of a periodic table. Helium (HE) 1S² NEON (NE) 1S²2S²2P⁶ ARGON (AR) 1S²2S²2P⁶3S²3P⁶ KRYPTON (KR) 1S²2S²2P⁶3S²3P⁶4S²4P⁶ © Pearson Education, Inc., or its affiliates. Carbon (C) 1S²2S²2P² SILICON (SI) 1S²2S²2P⁶3S²3P² Germanium (GE) 1S²2S²2P⁶3S²3P⁶4S²4P² Copyright © Pearson Education, Inc., or its affiliates. 11 Electron configurations in groups Representative elements Some elements of these groups are metals, some are not metals and some are metalloids. 19 Calculate Apply the concepts to this problem Example of problem 6.1 Calculate Apply the concepts to this problem. Copper, silver, gold and iron are transition metals. n. 5 Electron configurations in groups How can the elements be classified according to electron configurations? The 's' is in a sub-level s. Example: Copyright Uranium A © Pearson Education, Inc., or its affiliates. What does the group tell you? All rights reserved. 2 Reading the Table Periodic Table Symbolic Colors 3 Lidos A A Liquids Negros A A Blue Gases A A Red Gray A A Not found in nature *Note* The colors are specific to the periodic table in your textbook. 20 What is the period in which there is an element that tells us about the configuration? A electron A A that element? 2 Use Figure 6.13 to determine the configuration? the 's'. 3 Reading the Periodic Table The background colors of the squares are used to distinguish groups of elements in the periodic table. 13 Electron configurations in groups Representative elements In the 4A group the atom elements carbon, silicon and germanium, there are four electrons at the highest occupied energy level. 9 Electron configurations in groups Noble gases See the description of the occupied energy level higher for each element, which is highlighted in yellow. The elements of Group 2A are called earth alkaline metals. The name halogens comes from the combination of the Greek words, which means 'salt' and the Latin word 'genesis', which means 'to be born'. Copyright © Pearson Education, Inc., or its affiliates. 18 USING ENERGY SUBLEVELS TO WRITE ELECTRONIC CONFIGURATIONS Use use 6.9 (on pages 168 and 169 in your textbook) and Figure 6.13 on page 172 in your textbook to write the electron configuration of the nickel (NI). Elements in groups B are called transition elements. n. 7 Electron configurations in groups Noble Gases Helium, neon, argon and krypton are examples of noble gases, the elements in group 8a of the periodic table. The inner transition metals are characterized by the presence of electrons in the orbitals f. The s and p sublevels are completely filled with electrons, two electrons in the sublevel of s and six electrons in the sublevel of p. 4 Read the table Periodic Table The non-metals of group 7a are called halogens. Lithium (LI) 1S²2S¹ SODIUM (NA) 1S²2S²2P⁶3S¹ POTASSIUM (K) 1S²2S²2P⁶3S²3P⁶4S¹ Copyright © Pearson Education, Inc., or its affiliates. Most of them are solids, but some are gases at room temperature, and one, bromine, is a liquid. 22 End of 6.2 Copyright © Pearson Education, Inc., or its affiliates. In the volumes of these elements of group 1a, there is only one 's' at the highest occupied energy level. Put it all together: 1S²2S²2P⁶3S²3P⁶4S². 10 Electron configurations on the representative elements of Group 7 This figure shows a part of the periodic table containing groups 1a to 7a. 17 Electron configurations in groups Element blocks The periodic table is divided into sections, or blocks, that correspond to the highest sublevels occupied. These are sometimes called inert gases because they rarely participate in a reaction. n. These elements are characterized by the presence of electrons in the orbitals d. The noble gases neon and argon produce the colors in this sign of neon. Copyright © Pearson Education, Inc., or its affiliates. 6 Electron configurations in the groups can be classified into noble gases, representative elements, transition metals or transition elements based on their electron configurations. In Nickel, the first three energy levels are. they are. 6p3s3p2s2s2s1 6p3s3p2s2s2s1 HTTW Snteb NARIARUGIPNOC EHT OS

26/07/2020 - Page 1 of 6: Classifying parts of the world. The level of development, of a country shows how economically, socially, culturally or technologically advanced it is. 6 Fe²⁺ + Cr²⁺ O₇²⁻ + 14 H⁺ + 6 Fe³⁺ + 2 Cr³⁺ + 7 H₂O 6 Fe²⁺ + Cr²⁺ O₇²⁻ + 14 H⁺ + 6 Fe³⁺ + 2 Cr³⁺ + 7 H₂O final check of atom and charge balance confirms the equation is balanced. 01/01/2015 - Classifying disorders using the four ds of diagnoses. Paper Type: Free ... fantasies or behaviors existing over at least 6 months and directed at children 13 years of age or ... it also features elements of an anxiety disorder. The distress of the preoccupation of the disorder persists in spite of medical evaluations and

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